

# Co-ordination Compounds 1.

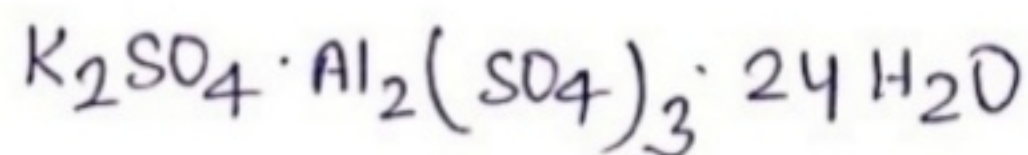
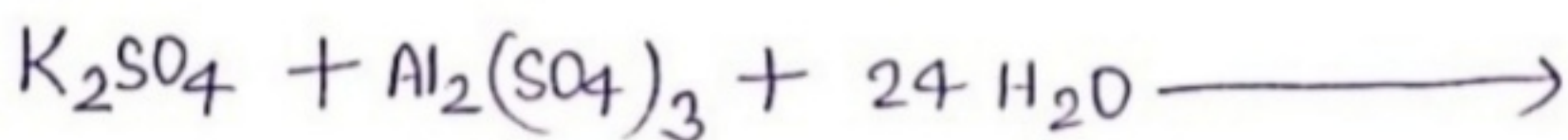
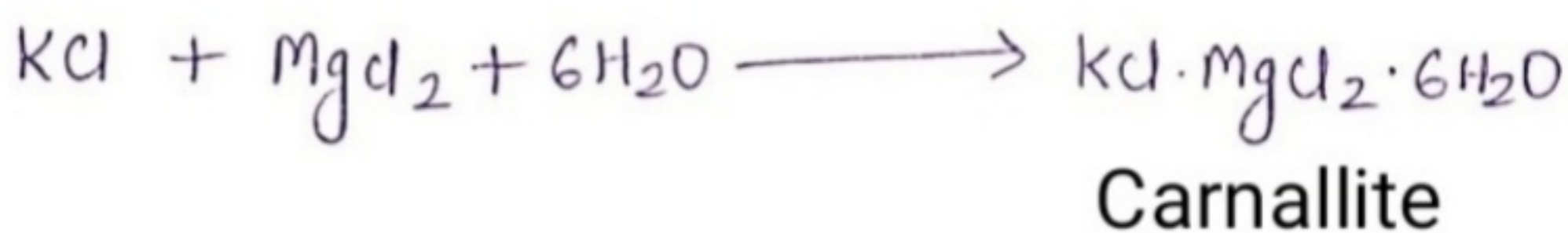
Degree-II (H) ,Paper-III ,Group-B

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## Double Salts and Complex Salts

\* Addition compounds are formed when stoichiometric amounts of two or more stable compounds join together.

For example :



Pot. alum



(Revision Notes)

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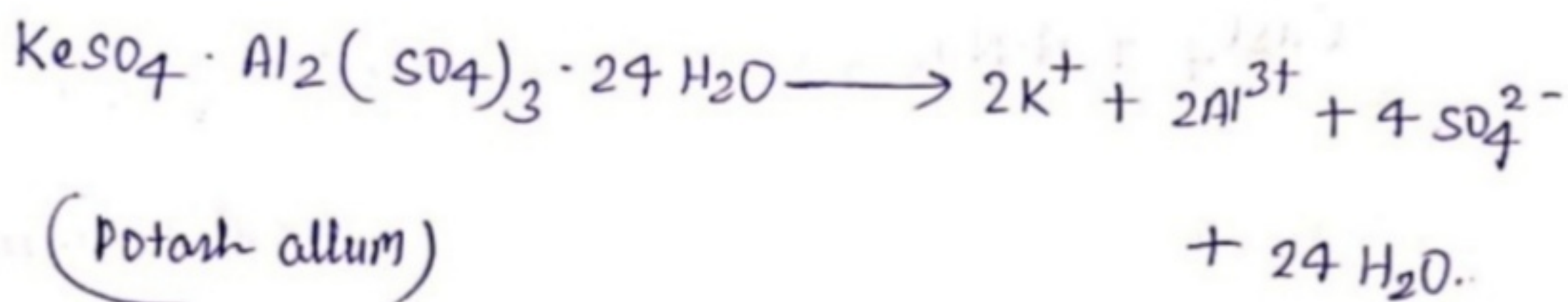
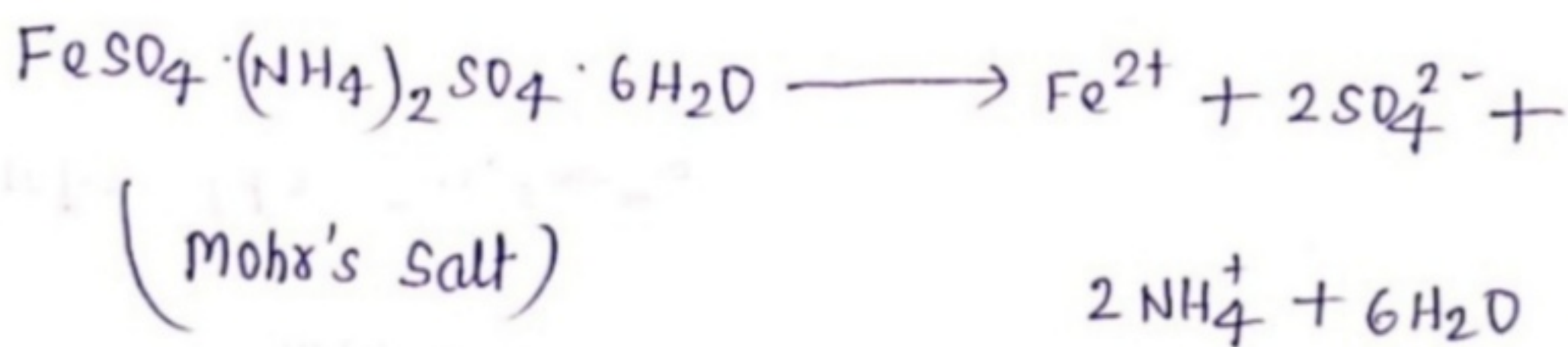
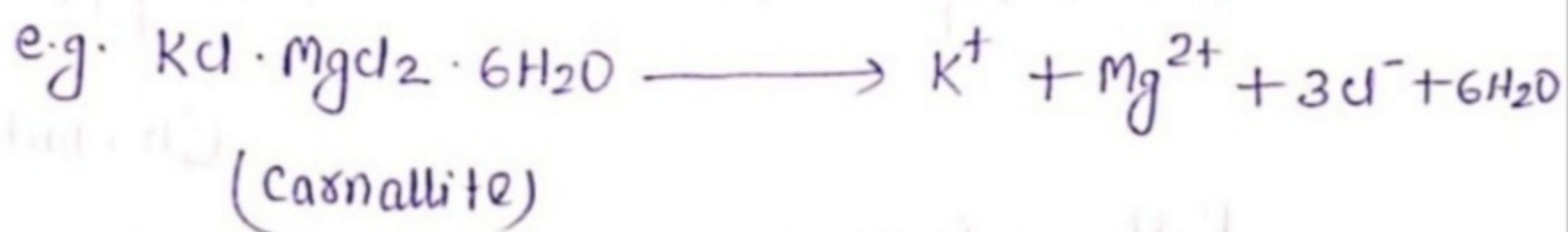
## Addition Compounds are of two 2.

### types :-

1. Those which lose their identity in solution:  
(Double Salts)
2. Those which retain their identity in solution:  
(Complex Salts)

## Double Salts

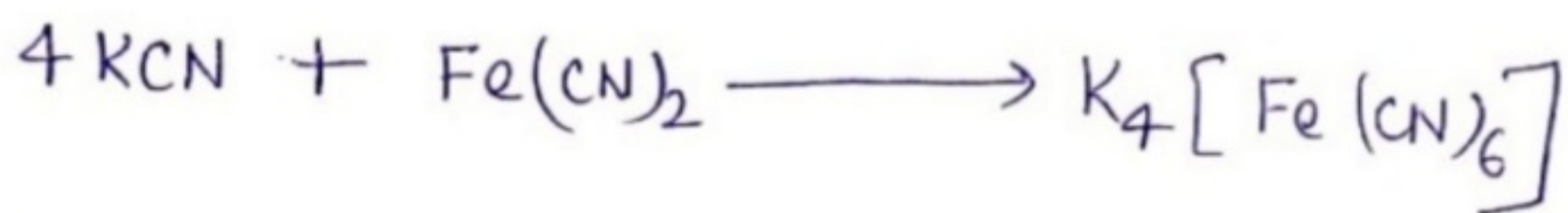
\* Double Salts are those addition compounds which exist only in crystal lattice but lose their identity in solution. Double salts ionise when dissolved in water.



# Complex Salts

\* Complex Salts are those addition compounds which retain their identity in solid / crystal lattice as well as in the solution.

e.g. Potassium ferrocyanide is a complex compound which is formed by adding KCN to a saturated solution of ferrous cyanide.



\*  $\text{K}_4[\text{Fe}(\text{CN})_6]$  is dissolved in water, the resulting solution does not give positive tests for ferrous or cyanide ions but we get a positive test for

